

THE EQUATION SHEET

Constants:

Avogadro's Number (N_A)	6.02×10^{23}
Universal Gas Constant (R)	8.314 J/mol·K or 0.0821 L·atm/mol·K
Planck's constant (h)	6.626×10^{-34} J·s
Rydberg Constant (R_H)	2.18×10^{-18} J
Speed of Light (c)	3.00×10^8 m/s
Charge of an Electron (q)	1.602×10^{-19}
Boltzmann Constant (k _B)	1.381×10^{-23} J/K
Molar Volume (V_{mol})	22.7 L/mol
Mass of Earth	5.97×10^{24} kg
Specific Heat Capacity of Water (C)	4.18 J/gK or 4.18 kJ/kgK
Ionic Product Constant of Water (K_w)	1.00×10^{-14} (mol/L) ² at 298 K (25°C)
Faraday's constant (F)	96 500 C/mol
STP conditions	273 K and 100 kPa

Basic Equations:

$n = \frac{m}{M_R}$	$n = cV$	$PV = nRT$
Order of reaction = m + n		$c_1V_1 = c_2V_2$
$n_{gas} = \frac{V}{22.7 \text{ mol/L}}$	$K_{SP} = K_c$ (Aqueous)	
% atom economy = $\frac{\text{molar mass of desired product}}{\text{molar mass of all reactants}} \times 100\%$		
Conversion factors:		
1 atm = 100 kPa	1 atm = 760 torr = 760 mm Hg	
1 nm = 10^{-9} m	0°C = 273.15 K	
1 dm ³ = 1 L = 1×10^{-3} m ³ = 1×10^3 cm ³ = 1×10^3 mL		
1 amu = 1.66×10^{-27} kg		

<p>Acid-Base Chemistry:</p> $pH = -\log[H_3O^+]$ $[H_3O^+] = 10^{-pH}$ $K_w = K_a \times K_b$ $pK_a + pK_b = pK_w$ $pK_a = -\log K_a$ $pK_b = -\log K_b$ $pK_b = 14 - pK_a$ $pH + pOH = 14$ $pOH = -\log[OH^-]$ $[OH^-] = 10^{-pOH}$ $pH_{Buffer} = pK_a - \log\left(\frac{[HA]}{[A^-]}\right)$	<p>Thermodynamics:</p> $\Delta H_{rxn} = H_P - H_R$ $q = \Delta H \text{ at constant pressure}$ $\Delta H^\circ = \frac{-Q}{\# \text{ mol}}$ $M_{Enthalpy} = \sum (E_k + E_p)$ $E_k = \frac{1}{2}mv^2$ $C = \frac{Q}{\Delta T}$ $Q = mc\Delta T$ $\Delta H_{rxn}^\circ = \sum [\Delta H_{f(P)}^\circ] - \sum [\Delta H_{f(R)}^\circ]$ $\Delta H_{rxn}^\circ = \sum D(\text{broken}) - \sum D(\text{formed})$ $\Delta S = k \ln W = \frac{q}{T} = \frac{\Delta H}{T} = S_{System} + S_{Surrounding}$ $\Delta S_{rxn}^\circ = \sum S_{(P)}^\circ - \sum S_{(R)}^\circ$ $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ $\Delta G_{rxn}^\circ = \sum \Delta G_{(P)}^\circ - \sum \Delta G_{(R)}^\circ$	<p>Chemical Kinetics & Equilibrium:</p> $Rate_{Reaction} = \frac{\Delta c}{\Delta t}$ $Rate_{Reaction} = k[A]^m[B]^n$ $E_A = -RT \ln\left(\frac{k}{A}\right)$ $t_{1/2} = \frac{0.693}{k}$ $t_{1/2} = \frac{1}{k[A]_0}$ $k = Ae^{-E_a/RT}$ $[A]_t = -kt + [A]_0$ $\ln[A]_t = -kt + \ln[A]_0$ $\ln \frac{k_1}{k_2} = \frac{E_a}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$ $K_C = \frac{[\text{Products}]^{nB}}{[\text{Reactants}]^{nA}}$ $K_P = K_C(RT)^{\Delta n}$ $\Delta G^\circ = -RT \ln K$	<p>Quantum Mechanics:</p> $\Delta E = \frac{hc}{\lambda}$ $c = \lambda \nu$ $\Delta E = R_H \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$ $E = hf$ $n? = 2d \sin \theta$
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<p>Nuclear Chemistry:</p> $E = mc^2$ ${}_{92}^{238}\text{U} \rightarrow {}_{90}^{234}\text{Th} + {}_2^4\text{He}$ ${}_0^1n \rightarrow {}_1^1H + {}_{-1}^0e$	<p>Gas:</p> $PV = nRT \quad \frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \frac{Rate_1}{Rate_2} = \sqrt{\frac{M_2}{M_1}}$ <p>STP conditions= 273 K and 100 kPa SATP conditions= 298 K and 100 kPa</p> $n_{gas} = \frac{V}{22.7 \text{ mol/L}}$	<p>Redox:</p> <p>Charge = Current × Time</p> $E_{cell}^\circ = E_{cathode}^\circ - E_{anode}^\circ$ $\Delta G^\circ = -nFE^\circ$
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Extras

<p>Solubility:</p> <table border="1" style="width: 100%;"> <tr> <th>Q_c</th> <th>K_c</th> <th>Q</th> <th>K_{sp} (Precipitate)</th> </tr> <tr> <td><</td> <td>Prod Fav</td> <td><</td> <td>No</td> </tr> <tr> <td>=</td> <td>EQ</td> <td>=</td> <td>No</td> </tr> <tr> <td>></td> <td>React Fav</td> <td>></td> <td>Yes (Super Saturated)</td> </tr> </table> <p>Aufbau Principle: Build up electrons one by one. 1K(2)2L(8)3M(18)4N(32)5O(50)6P(72)7Q(98)</p>	Q _c	K _c	Q	K _{sp} (Precipitate)	<	Prod Fav	<	No	=	EQ	=	No	>	React Fav	>	Yes (Super Saturated)	<p>Formations:</p> <ol style="list-style-type: none"> Acid + Metal = Salt + Hydrogen Gas Ex. $2\text{HCl}_{(aq)} + \text{Zn}_{(s)} \rightarrow \text{ZnCl}_{2(s)} + \text{H}_{2(g)}$ Acid + Base = Salt + Water Ex. $\text{HCl}_{(aq)} + \text{NaOH}_{(aq)} \rightarrow \text{NaCl}_{(s)} + \text{H}_2\text{O}_{(l)}$ Acid + Metal Carbonate = CO₂ + H₂O + Salt Ex. $\text{CaCO}_{3(s)} + \text{HCl}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)} + \text{CaCl}_{2(s)}$ Metal Oxide + Acid → Salt + Water Ex. $\text{MgO}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{MgCl}_{2(s)} + \text{H}_2\text{O}_{(l)}$
Q _c	K _c	Q	K _{sp} (Precipitate)														
<	Prod Fav	<	No														
=	EQ	=	No														
>	React Fav	>	Yes (Super Saturated)														

Periodic Table of Electronegativities

H 2.20																	He n.a.	
Li 0.98	Be 1.57											B 2.04	C 2.55	N 3.04	O 3.44	F 3.98	Ne n.a.	
Na 0.93	Mg 1.31											Al 1.61	Si 1.90	P 2.19	S 2.58	Cl 3.16	Ar n.a.	
K 0.82	Ca 1.00	Sc 1.36	Ti 1.54	V 1.63	Cr 1.66	Mn 1.55	Fe 1.83	Co 1.88	Ni 1.91	Cu 1.90	Zn 1.65	Ga 1.81	Ge 2.01	As 2.18	Se 2.55	Br 2.96	Kr 3.00	
Rb 0.82	Sr 0.95	Y 1.22	Zr 1.33	Nb 1.60	Mo 2.16	Tc 1.90	Ru 2.20	Rh 2.28	Pd 2.20	Ag 1.93	Cd 1.69	In 1.78	Sn 1.96	Sb 2.05	Te 2.10	I 2.66	Xe 2.60	
Cs 0.79	Ba 0.89	La 1.10	Hf 1.30	Ta 1.50	W 2.36	Re 1.90	Os 2.20	Ir 2.20	Pt 2.28	Au 2.54	Hg 2.00	Tl 1.62	Pb 2.33	Bi 2.02	Po 2.00	At 2.20	Rn n.a.	
Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Uub		Tl	Pb	Bi	Po	At	Rn

Polyatomic Ions:

Acetate	CH ₃ COO ⁻ or C ₂ H ₃ O ₂ ⁻	Hydroxide	OH ⁻
Aluminate	AlO ₂ ⁻ , Al ₂ O ₄ ²⁻	Hypobromite	BrO ⁻
Amide	NH ₂ ⁻	Hypochlorite	ClO ⁻
Ammonium	NH ₄ ⁺	Hypoiodite	IO ⁻
Antimonate	SbO ₄ ³⁻	Hypophosphite	PO ₂ ³⁻
Antimonite	SbO ₃ ³⁻	Hyposulfite	SO ₂ ²⁻
Arsenate	AsO ₄ ³⁻	Iodate	IO ₃ ⁻
Arsenite	AsO ₃ ³⁻	Iodite	IO ₂ ⁻
Bicarbonate (hydrogen carbonate)	HCO ₃ ⁻	Manganate	MnO ₄ ²⁻
Bromate	BrO ₃ ⁻	Nitrate	NO ₃ ⁻
Bromite	BrO ₂ ⁻	Nitrite	NO ₂ ⁻
Carbide	C ₂ ²⁻	Oxalate	C ₂ O ₄ ²⁻
Carbonate	CO ₃ ²⁻	Ozonide	O ₃ ⁻
Chlorate	ClO ₃ ⁻	Perbromate	BrO ₄ ⁻
Chlorite	ClO ₂ ⁻	Perchlorate	ClO ₄ ⁻
Chromate	CrO ₄ ²⁻	Periodate	IO ₄ ⁻
Chromite	CrO ₂ ⁻	Permanganate	MnO ₄ ⁻
Cyanate	OCN ⁻	Peroxide	O ₂ ²⁻
Cyanide	CN ⁻	Phosphate	PO ₄ ³⁻
Dichromate	Cr ₂ O ₇ ²⁻	Phosphite	PO ₃ ³⁻
Dihydrogen arsenate	H ₂ AsO ₄ ⁻	Plumbate	PbO ₃ ²⁻
Dihydrogen phosphate	H ₂ PO ₄ ⁻	Plumbite	PbO ₂ ²⁻
Dihydrogen phosphite	H ₂ PO ₃ ⁻	Stannate	SnO ₃ ²⁻
Disulfide	S ₂ ²⁻	Stannite	SnO ₂ ²⁻
Ferrate	FeO ₄ ²⁻	Sulfate	SO ₄ ²⁻
Hydrogen carbonate (bicarbonate)	HCO ₃ ⁻	Sulfite	SO ₃ ²⁻
Hydrogen arsenate	HAsO ₄ ²⁻	Superoxide	O ₂ ⁻
Hydrogen phosphate	HPO ₄ ²⁻	Tartrate	(CH(OH)COO) ₂ ²⁻
Hydrogen phosphite	HPO ₃ ²⁻	Tellurate	TeO ₄ ²⁻
Hydrogen sulfate	HSO ₄ ⁻	Tellurite	TeO ₃ ²⁻
Hydrogen sulfite	HSO ₃ ⁻	Thiocyanate	SCN ⁻
Hydronium	H ₃ O ⁺	Thiosulfate	S ₂ O ₃ ²⁻

Periodic Table of the Elements

1 IA 1A																	18 VIIIA 8A			
1 H Hydrogen 1.008	2 IIA 2A												3 B Boron 10.811	4 C Carbon 12.011	5 N Nitrogen 14.007	6 O Oxygen 15.999	7 F Fluorine 18.998	8 Ne Neon 20.180		
3 Li Lithium 6.941	4 Be Beryllium 9.012	3 IIIB 3B	4 IVB 4B	5 VB 5B	6 VIB 6B	7 VIIB 7B	8 VIII 8		9 VIII 9		10 VIII 10		11 IB 1B	12 IIB 2B	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.065	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.88	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.64	33 As Arsenic 74.922	34 Se Selenium 78.972	35 Br Bromine 79.904	36 Kr Krypton 83.80			
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.94	43 Tc Technetium 98.906	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.905	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.71	51 Sb Antimony 121.757	52 Te Tellurium 127.6	53 I Iodine 126.905	54 Xe Xenon 131.29			
55 Cs Cesium 132.905	56 Ba Barium 137.327	57-71 Lanthanide Series		72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.967	80 Hg Mercury 200.59	81 Tl Thallium 204.384	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [209]	85 At Astatine [210]	86 Rn Radon 222.018		
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103 Actinide Series		104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [263]	107 Bh Bohrium [264]	108 Hs Hassium [265]	109 Mt Meitnerium [266]	110 Ds Darmstadtium [268]	111 Rg Roentgenium [269]	112 Cn Copernicium [271]	113 Uut Ununtrium [272]	114 Fl Flerovium [274]	115 Uup Ununpentium [275]	116 Lv Livermorium [276]	117 Uus Ununseptium [278]	118 Uuo Ununoctium [279]		
Lanthanide Series		57 La Lanthanum 138.905	58 Ce Cerium 140.12	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.24	61 Pm Promethium [144.913]	62 Sm Samarium 150.36	63 Eu Europium 151.965	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.50	67 Ho Holmium 164.930	68 Er Erbium 167.26	69 Tm Thulium 168.934	70 Yb Ytterbium 173.054	71 Lu Lutetium 174.967				
Actinide Series		89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.063	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.067	98 Cf Californium 251.080	99 Es Einsteinium [252]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]	103 Lr Lawrencium [260]				